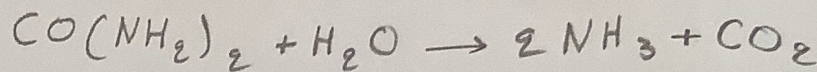


Q5:



$$m = 2000 \text{ g}$$

$$T = 294 \text{ K}$$

$$p = 101,5 \text{ kPa}$$

$$M_m = 60$$

$$n = 33,3$$

↓
↓

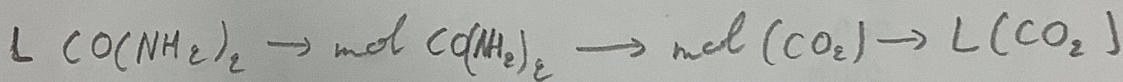
$$n_{\text{CO}(\text{NH}_2)_2} = \frac{2000 \text{ g}}{60 \text{ g/mol}} = 33,3 \text{ mol}$$

d'abord on calcule le volume de $\text{CO}(\text{NH}_2)_2$

$$V = \frac{mRT}{MP} = \frac{nRT}{P} = \frac{33,3 \text{ mol} \cdot 8,31 \frac{\text{L kPa}}{\text{K mol}} \cdot 294 \text{ K}}{101,5 \text{ kPa}}$$

$$V = 801,5 \text{ L}$$

Donc:



$$\frac{801,5}{60} \times 44 = 587,76 \text{ L}$$